



Cambridge IGCSE™

COMBINED SCIENCE

0653/32

Paper 3 Theory (Core)

May/June 2023

MARK SCHEME

Maximum Mark: 80

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the May/June 2023 series for most Cambridge IGCSE, Cambridge International A and AS Level and Cambridge Pre-U components, and some Cambridge O Level components.

This document consists of **12** printed pages.

Generic Marking Principles

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

- | | |
|---|--|
| 1 | Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly. |
| 2 | The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored. |
| 3 | Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection). |
| 4 | The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted. |
| 5 | <p><u>'List rule' guidance</u></p> <p>For questions that require <i>n</i> responses (e.g. State two reasons ...):</p> <ul style="list-style-type: none"> • The response should be read as continuous prose, even when numbered answer spaces are provided. • Any response marked <i>ignore</i> in the mark scheme should not count towards <i>n</i>. • Incorrect responses should not be awarded credit but will still count towards <i>n</i>. • Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should not be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response. • Non-contradictory responses after the first <i>n</i> responses may be ignored even if they include incorrect science. |

6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

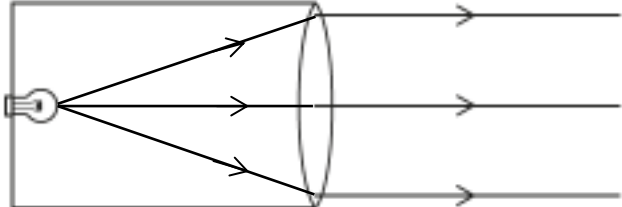
Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

Question	Answer	Marks												
1(a)(i)	D ; C ;	2												
1(a)(ii)	aorta (circled) ;	1												
1(b)(i)	star jumps ;	1												
1(b)(ii)	(running) 110 (bpm) and (step-ups) 96 (bpm) ; (difference =) 14 (bpm) ;	2												
1(c)	<table border="1"> <thead> <tr> <th>name of part</th><th>letter in Fig. 1.3</th><th>function</th></tr> </thead> <tbody> <tr> <td>red blood cell</td><td>P</td><td>transport of oxygen</td></tr> <tr> <td>plasma</td><td>S</td><td>transport of hormones</td></tr> <tr> <td>white blood cells</td><td>Q</td><td>phagocytosis / produce antibodies</td></tr> </tbody> </table> one mark for each correct row	name of part	letter in Fig. 1.3	function	red blood cell	P	transport of oxygen	plasma	S	transport of hormones	white blood cells	Q	phagocytosis / produce antibodies	3
name of part	letter in Fig. 1.3	function												
red blood cell	P	transport of oxygen												
plasma	S	transport of hormones												
white blood cells	Q	phagocytosis / produce antibodies												

Question	Answer	Marks
2(a)(i)	(concentrated) high ratio of, solute / sodium chloride / ions, to solvent / water OR large (relative) amount of, solute / sodium chloride / ions ; (aqueous) (dissolved) in water / the solvent is water ;	2
2(a)(ii)	chlorine (gas) ;	1
2(a)(iii)	<u>ionic</u> ;	1
2(b)	white, precipitate / solid ;	1

Question	Answer	Marks
2(c)(i)	sodium hydroxide / NaOH / sodium carbonate / Na_2CO_3 / sodium hydrogen carbonate / NaHCO_3 ;	1
2(c)(ii)	increases / rises / goes up to 7 ;	1
2(c)(iii)	exothermic ;	1
2(d)	(electrons) 11 ; (neutrons) 12 ;	2

Question	Answer	Marks
3(a)(i)	10 (s) ;	1
3(a)(ii)	20 (s) ;	1
3(a)(iii)	X on line between t = 10 s and t = 25 s ;	1
3(b)	90 (km/h) ;	1
3(c)	speed = distance $\div$ time (<i>in any form</i>) / distance = 30×65 ; 1950 (m) ;	2
3(d)(i)	refraction ; speed ;	2
3(d)(ii)	 ;	1

Question	Answer	Marks
4(a)	test-tube B : (more) carbon dioxide used up in photosynthesis (than produced in respiration) ; test-tube C : no photosynthesis / only respiration ; carbon dioxide made in respiration ;	3
4(b)(i)	support ;	1
4(b)(ii)	guard ; stomata ; water ;	3
4(b)(iii)	decreased humidity / increased temperature ;	1
4(b)(iv)	xylem ;	1

Question	Answer	Marks
5(a)(i)	(physical property) high density / high melting point / electrical conductor / heat or thermal conductor ; (chemical property) low reactivity ;	2
5(a)(ii)	stronger / harder ;	1
5(b)(i)	$..2..CuO + ..(1)..C \rightarrow ..(1)..CO_2 + ..2..Cu$;	1
5(b)(ii)	(carbon is) oxidised AND (reason) it gains oxygen ;	1
5(b)(iii)	aluminium is more reactive (than carbon) ;	1

Question	Answer	Marks
5(b)(iv)	<i>any two from:</i> noble gases / argon, are, unreactive / inert ; (because) they have full outer electron shells ; do not lose / gain, electrons ;	2

Question	Answer	Marks
6(a)	<i>any two from:</i> hydroelectric (behind dam) ; waves ; tides ;	2
6(b)(i)	235 (°C) ;	1
6(b)(ii)	geothermal ;	1
6(b)(iii)	<i>any three from:</i> water turns to steam ; at high pressure ; drives / moves, turbine ; (which) turns generator ; kinetic energy is converted to electrical energy ;	3
6(c)(i)	chemical (potential) ;	1
6(c)(ii)	gravitational potential ; longer / more time ;	2

Question	Answer	Marks
7(a)(i)	arrow from anchovy to bluefish AND from shrimp to bluefish ;	1
7(a)(ii)	shrimp / mussel ;	1

Question	Answer	Marks
7(a)(iii)	the Sun ;	1
7(b)	any two from: genetically identical offspring ; one parent / does not need two parents / does not need male and female parents ; no fertilisation / no fusion of nuclei / no fusion of gametes / no gametes involved ;	2
7(c)(i)	for insulation / energy store / energy source ;	1
7(c)(ii)	ethanol ;	1
7(c)(iii)	(gap one) large AND (gap two) small ; enzymes ;	2

Question	Answer	Marks
8(a)(i)	methane + oxygen $\rightarrow$ carbon dioxide + water ;	1
8(a)(ii)	both correct bonding pairs ; molecule fully correct, i.e. four additional electrons on O and no extra electrons ;	2
8(b)(i)	contains a double bond ;	1
8(b)(ii)	(aqueous) bromine / Br ₂ ;	1
8(b)(iii)	addition ;	1
8(c)	any two from: separation: gas particles are, far apart / separated / not touching / spread out / liquid particles are, close together / touching / not spread out / not separated ; movement: gas particles, move fast(er) / move freely / move in all directions / move randomly / liquid particles, move (more) slowly / move over each other ; forces: gas particles have, no / weak forces of attraction between particles / liquids particles have stronger forces of attraction ;	2

Question	Answer	Marks
9(a)	resistance = voltage $\div$ current (<i>in any form</i>) / ($R = 9 \div 1.5$; ($R =$) 6 ; ohms / Ω ;	3
9(b)(i)	<u>variable</u> resistor ;	1
9(b)(ii)	current decreases ;	1

Question	Answer	Marks
9(b)(iii)	amplitude ;	1
9(c)	lamp symbol ; connected in parallel above or below loudspeaker, but below switch ;	2